Determining limiting reactant worksheet

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Determine the mass of producable iodine I2? b) If only 0.160 mole iodine is produced in the above case, I2 is produced. i) What iodine mass was produced? ii) What percentage of iodine yield was produced. 2. Reaction to create zinc sulfur according to zinc and sulfur equation. Zn + S ---------> If 25.0 g of zinc and 30.0 g of sulfur are mixed, a) What chemical is the limiting reactive? b) How many grams of ZnS will occur? c) How many grams will remain from the oversead after the reaction is over? 3. What element is more when 3.00 grams of Mg of pure oxygen is fired at 2.20 grams? What mass is lost? How much is the mass of MGO formed? 4. How many grams of Al2S3 occur when heated to 5.00 grams Al 10.0 grams S? 5. When MoO3 and Zn are heated together, what ZNO mass occurs when 3 Zn(s) + 2 MoO3(s) ----------> Mo2O3(s) + 3 ZnOs(s) react with 20.0 grams of MoO3? 6. Silver nitrate, AgNO3, react with ferric chloride, FeCl3, give silver chloride, AgCl, and ferric nitrate, Fe(NO3)3. In a specific experiment, it was planned to mix a solution containing 25.0 g of AgNO3 with another solution containing 45.0 grams of FeCl3. a) Write the chemical equation for the reaction. b) Which of the limiting reacters? c) What is the highest number of moles of AgCl, which can be obtained from this mixture? d) What is the maximum number of grams the OSCE can receive? e) How many grams of reaktan will it remain after the reaction is over? 7. Solid calcium carbonate, CaCO3, is able to remove sulfur dioxide from waste gases by reaction: CaCO3 + SO2 + caso3 ------> other reacters + other products In a specific experiment, CaCO3 was exposed in the presence of an excessive amount of other chemicals necessary for the 255 g SO2 135 g reaction. a) What is the theoretical efficiency of CaSO3? b) If only 198 g of CaSO3 was isolated from products, how much was the precentage yield of CaSO3 in this experiment? 8. A research consultant told a chemist to make 100 g of chlorobenzene with chlorine reaction chlorine and expect a higher yield of 65%. If the yield is 65%, what is the minimum amount of benzene that can give 100 g of chloroobenzene? Reaction flour equation: C6H6 + Cl2 ----------> C6H5Cl + HCl benzene chloroobenzen 9. Some salts of benzoic acid have been used as food additives for years. Benzoic acid can be done by potassium salt, potassium benzoate, potassium permanganate action on toluene. C7H8 + 2 KMnO4 -------> KC7H5O2 + 2 MnO2 + KOH + H2O toluene potassium benzoate potassium benzoate if potassium benzoate yield cannot realistically be expected to be more than 68%, what is the minimum gram required to achieve this efficiency when producing the minimum number of toluene kc7H5O2 10.0 g? 10. Aluminum is soluble in a NaOH aqueous solution according to the following reaction: 2 NaOH + 2 Al + 2 H2O -----> 2 NaAlO2 + 3 H2 If NaOH 84.1 g and Al 51.0 g response: i) Which is the limiting reactive? ii) How much of the other reagent remains? iii) In what mass is hydrogen mass produced? 11. Dimethylhydrazine, (CH3)2NNH2 was used as fuel for the Apollo Lunar Landing Module, N2O4 was used as oxidant. Products There are H2O, N2 and CO2. i) Write a balanced chemical equation for combustion reaction. ii) If 150 kg (CH3)2NNH2 reacts with 460 kg N2O4, what is the theoretical efficiency of N2? iii) If the yield of 30 kg of N2 gas represents a yield of 68%, what mass of N2O4 was used in the reaction? 12. Herdative reaction with oxygen to give magnesium metal magnesium oxide, MgO. If 5.00 g mg and 5.00 g of O2 are allowed to react, how much does the weight of MgO occur and what weight is oversted? 13. Adypic acid, C6H10O4, is a raw material for nylon making and can be prepared in the laboratory with the following reaction between cycloheksen, C6H10 and sodium dichromate, Na2Cr2O7 sulphuric acid. There are 3 C6H10(l) + 4 Na2Cr2O7(aq) + 16 H2SO4(aq) ---------> 3 C6H10O4(aq) + 4 Cr2(SO4)3(aq) + 4 Na2SO4(aq) + 16 H2O Side reactions. This plus reduces losses during the purification of the product, reducing overall efficiency. The typical yield of purified adipic acid is 68.6%. (a) How many grams of cyclohecs does the 68.6% yield require to prepare 12.5 grams of adyspic acid? (b) The only available source of sodium dichromate is Na2Cr2O7.2H2O, which is dihydrate. (Since the reaction occurs in an aqueous environment, dihydrate water causes no problems, but this contributes to the mass taken from the reacter). How many grams of this dihydrate are required in the preparation of 12.5 grams of adyspic acid of 68.6%? Responses Answers